

Erratum

Erratum to “On the use of ammonia electrolysis for hydrogen production”  
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The publisher regrets that there were errors in the figures published in the above-mentioned article.

The errors were as follows:

1. The image for Fig. 1 should have been the image used in Fig. 4.
2. The image for Fig. 2 should have been the image used in Fig. 1.

3. The image for Fig. 3 should have been the image used in Fig. 2.
4. The image for Fig. 4 should have been the image used in Fig. 3.

Please see the corrected figures reproduced here.

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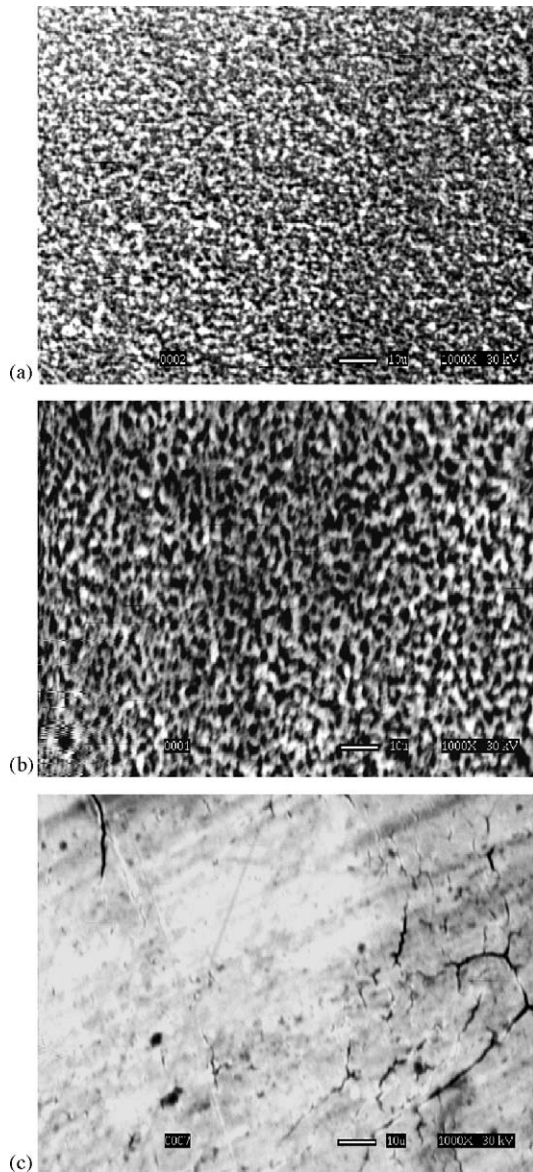


Fig. 1. Scanning electron photomicrographs. Magnification 1000×, voltage 30 kV: (a) on Pt black electrode, a rough deposit is observed; (b) on Pt–Ru electrode, rough deposit is observed as well; (c) Pt–Ir with a smoother, denser morphology.

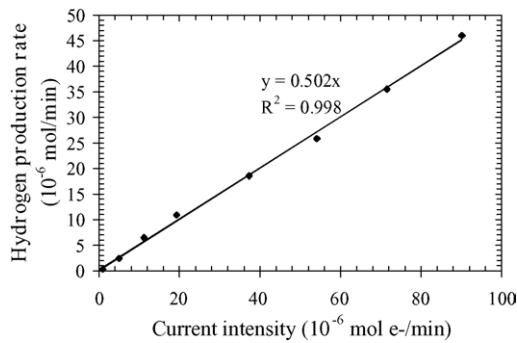


Fig. 2. Hydrogen production rate showing that the reaction represented by Eq. (2) boasts a 100% faradaic efficiency (slope is 0.5).

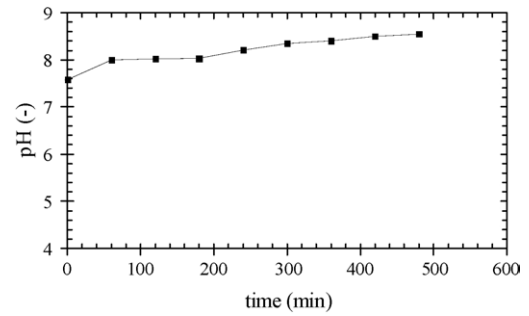


Fig. 3. pH of the bubbling water with respect to time as produced hydrogen is bubbled through. pH is used to calculate the concentration of ammonia in the liquid and gaseous phase under the assumption of vapor–liquid equilibrium.

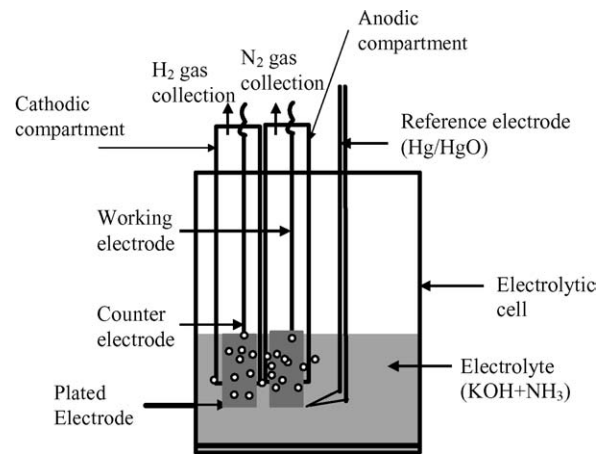


Fig. 4. Three-compartment electrolytic cell for the electrolysis of ammonia. All the experiments reported in the paper were performed in this cell. The electrodes are immersed in the solution.